

# Metals and their Compounds

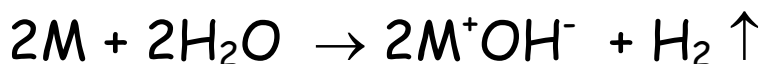
## Lecture 1.4

Active metals are Groups 1, 2 and 3 (aluminium).  
Most reactive metals - in nature always occur as their IONS ( $\text{Na}^+$ ,  $\text{K}^+$   $\text{Mg}^{2+}$   $\text{Ca}^{2+}$  in sea water)

Group 1 metals as example :

Reactivity depends on size (lithium smallest, caesium largest)

Forms cations easily :



Name	Symbol	Reacts with water
Lithium	Li	smoothly
Sodium	Na	vigourously
Potassium	K	violently
Rubidium	Rb	detonates
Caesium	Cs	detonates

This shows that reactivity increases smoothly as we go down Periodic Column.