

Metals and their Compounds

Lecture 5.4

Transition metals (23.7 and Ch 24 BLB)

Formation of *coordination compounds*

CuSO_4 (copper(II) sulfate) solution contains the ion $\text{Cu}^{2+}_{\text{aq}}$ usually written as $[\text{Cu}(\text{H}_2\text{O})_6]^{2+}$

This is pale blue. Reactions demonstrated have formed two new *coordination compounds* of Cu^{2+} (these are *not* a redox reactions)

